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3.  $K = \frac{[N_2]^{1/2} [O_2]^{1/2}}{[NO]^{1/2} [H_2O]^{1/2}}$  Use of the Lewis and Randall rule gives:

$[NO] = 2f_{NH_3} \approx 2 \times 1 \text{ atm}$ .

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BASICS 2.1

ARRHENIUS

EQUATION The main

assumption behind the

Arrhenius expression is

that  $k = A e^{-E_a/RT}$  (

).

This is an approximation, but it



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works quite well. The rate coefficient is the term that is a function of temperature but may also depend

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Chemical Equilibrium  
 $dG = - SdT + VdP + \sum_{j=1}^n \mu_j dn_j$   
 $\mu_j$ : chemical potential for species j.  $G = G(T, P)$

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$$\mu_j = G_j + RT \ln a_j, \quad a_j = f_j / f_j^{\text{standard}}$$
$$G_i = -RT \ln K_i$$
Standard state: pure species  $j$  at 1 atm and system temperature.

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